

## Kinetics and Mechanism of *meso*-Tetraphenylporphyriniron(III) Chloride Catalyzed Oxidation of Indole by Peroxomonosulphate

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(Received: 24 April 2012;

Accepted: 9 February 2013)

AJC-12952

Mechanistic study on *meso*-tetraphenylporphyriniron(III) chloride (TPP) catalyzed oxidation of indole by peroxomonosulphate in aqueous acetonitrile medium have been carried out. The reaction follows a fractional order (0.57) with respect to substrate and first order with respect to oxidant. The order with respect to catalyst was found to be one. Increase in percentage of acetonitrile decreased the rate of the reaction. The added  $H^+$  did not affect the reaction rate. The reaction fails to initiate polymerization and free radical mechanism is ruled out. Activation and thermodynamic parameters have been computed. Nucleophilic attack of the ethylenic bond on the persulphate oxygen in presence of *meso*-tetraphenylporphyriniron(III) chloride is envisaged to explain the reactivity. A suitable kinetic scheme based on the observations was proposed. Significant catalytic activity is observed for the reaction system in the presence of *meso*-tetraphenylporphyriniron(III) chloride.

**Key Words:** Indole, Peroxomonosulphate, *Meso*-tetraphenylporphyriniron(III) chloride, Kinetics, Oxidation.

### INTRODUCTION

Indole, an electron-rich heteroaromatic nitrogen compound, is found in many naturally occurring compounds such as various plant alkaloids and fungal metabolites<sup>1,2</sup> usually as metabolites of tryptophan. Indole, on oxidation gives products which are involved in significant biological processes<sup>3</sup>. Indoles possess significant biological activity<sup>4,5</sup>. Lawson and Witkop<sup>6</sup> as well as Finch and Taylor<sup>7</sup> have shown that N-bromosuccinimide can be used to convert indoles to oxindoles. Rangappa *et al.*<sup>8</sup> have examined the kinetics of oxidation of indole by chloramines-T in the presence of Os(VIII) in alkaline medium. Peroxomonosulphate can also be used to convert indoles to oxindoles<sup>9</sup>.

Peroxomonosulphate (PMS) ion,  $HSO_5^-$ , is an anionic oxidant which is usually represented as oxone,  $2KHSO_5 \cdot KHSO_4 \cdot K_2SO_4$ . Studies on the oxidation of organic compounds and inorganic complexes by PMS have been reported<sup>10</sup>. The synthetic utility of potassium peroxomonosulphate as an oxidant has been studied in detail<sup>11</sup>. It is a powerful oxidant and oxidizes alcohols, ketones, carboxylic acids, alkenes, arenes, phenols, amines and sulfides<sup>12,13</sup>. It hydroxylates alkanes and aromatic compounds<sup>14</sup>. The peroxy anion oxidation of indole has been carried out because its distinct reactivity is appropriate for an independent and precise kinetic study on the steps of the mechanism<sup>15</sup>.

Oxidation of organic compounds catalyzed by synthetic metalloporphyrin complexes has been one of the hot areas in

chemistry during the last decade<sup>16</sup>. The metalloporphyrins can catalyze the oxidation of a variety of organic substrates with many different terminal oxidants<sup>16,17</sup>. The catalytic properties of the transition metal porphyrins are due to the fact that an oxo-transition metal porphyrin intermediate is formed, which can transfer the oxygen atom to a substrate or can accept an electron from the substrate<sup>16</sup>. Many stable oxo-transition metal complexes are known and have been characterized, but several of these complexes are inert and non-reactive as oxygen transfer reagents. Groves and coworkers<sup>18</sup> described the use of *meso*-tetraphenylporphyriniron(III) chloride (TPP) in combination with the lipophilic iodosylbenzene, first used *in vivo* by Ullrich, for the epoxidation of olefins and the hydroxylation of alkanes.

Literature survey reveals that no report was available on the kinetics of oxidation of *meso*-tetraphenylporphyriniron(III) chloride catalyzed oxidation of indole, hence we have carried out this work with peroxomonosulphate.

### EXPERIMENTAL

All the chemicals and solvents used were of analytical grade (Merck, India). Indole, oxone and *meso*-tetraphenylporphyriniron(III) chloride (TPP) (Sigma Aldrich) were used as such. All the solutions used in the study were made by using doubly distilled water. All the reagents were prepared freshly and used in the reaction. All the reactions were carried out in a thermostat and the temperature was controlled to  $\pm 0.1$  °C.

**Kinetic measurements:** The kinetic studies were carried out in an aqueous acetonitrile medium. The reactions were performed by maintaining a large excess of [indole] over [PMS]. The reaction mixture was homogeneous throughout the course of the reaction. The reaction's progress was monitored for at least two half-lives by iodometric estimation of unchanged oxidant at regular time intervals. The rate constants ( $k_{\text{obs}}$ ) were evaluated from the slopes of linear plots of  $\log$  [titre] *versus* time.

**Stoichiometry:** Solutions of indole containing an excess of PMS were kept overnight at room temperature. Titrimetric estimation of the concentration of PMS consumed and assuming that all the indole taken had reacted, the stoichiometry of indole:PMS was found to be 1:1.

**Product analysis:** The reaction mixture after 48 h, was extracted with chloroform, dried over anhydrous  $\text{Na}_2\text{SO}_4$  and then evaporated. The residual brown solid obtained was subjected to column chromatography using silica gel and eluted with benzene and then evaporated. The melting point of the compound was found to be 116-118 °C confirms oxindole (literature value is 118-120 °C) as the main product of the reaction. The product was also confirmed by UV-visible absorption spectra at  $\lambda_{\text{max}}$  of 397 nm (Fig. 7) which is in agreement with literature values.

**Data analysis:** Correlation studies were carried out using Microcal origin (version 6) computer software. The goodness of the fit was discussed using the correlation coefficient,  $r$ , in the case of simple linear regression and  $R$  in the case of multiple linear regressions.

## RESULTS AND DISCUSSION

Factors influencing the rate of TPP catalyzed oxidation of indole by peroxomonosulphate such as [indole], [PMS], [TPP],  $[\text{H}^+]$  and dielectric constant have been studied. Rate and activation parameters were evaluated.

**Effect of [indole]:** At a constant [PMS], [TPP],  $[\text{H}^+]$  and fixed percentage of acetonitrile, kinetic runs were carried out with various initial concentrations of indole, which yielded rate constants whose values depended on [indole]. The pseudo-first order rate constants ( $k_{\text{obs}}$ ) thus obtained were found to increase with [indole] (Table-1) over a range of [indole] used ( $3.0\text{--}7.0 \times 10^{-2} \text{ mol dm}^{-3}$ ). The plot (Fig. 1) of  $\log k_{\text{obs}}$  *versus*  $\log$  [indole] is linear with a slope of 0.57 showing that the reaction is fractional order in [indole]. The plot (Fig. 2) of  $1/k_{\text{obs}}$  *versus*  $1/[\text{indole}]$  is linear with definite intercept on the rate ordinate, which indicates the operation of Michaelis-Menton mechanism, where  $K_3$  is the equilibrium constant for the formation of Michaelis-Menton complex and  $k_4$  the rate constant for the decomposition of Michaelis-Menton complex. Usually first step is a fast pre-equilibrium and the electron transfer step is rate determining<sup>19</sup>.

**Effect of [PMS]:** The order of the reaction with respect to PMS was to be strictly unity as shown by the linearity of  $\log$  [PMS] *versus* time plots, over 80 % of the reaction. Further, the reaction rate decreases with increase in the concentration of oxidant, [PMS] ( $1.5\text{--}3.5 \times 10^{-3} \text{ mol dm}^{-3}$ ) at fixed [indole], [TPP],  $[\text{H}^+]$  and percentage of acetonitrile (Table-1).

**Effect of [TPP]:** At a constant [indole], [PMS],  $[\text{H}^+]$  and fixed percentage of acetonitrile, kinetic runs were carried out

TABLE-1  
PSEUDO-FIRST ORDER RATE CONSTANTS FOR  
THE TPP CATALYSED OXIDATION OF INDOLE BY  
PEROXOMONOSULPHATE AT 301 K

[Indole] $\times 10^2$ (mol dm <sup>-3</sup> )	[PMS] $\times 10^3$ (mol dm <sup>-3</sup> )	[TPP] $\times 10^7$ (mol dm <sup>-3</sup> )	CH <sub>3</sub> CN % (V/V)	$k_{\text{obs}} \times 10^4$ (s <sup>-1</sup> )
3.0	2.5	4.0	80	1.52
4.0	2.5	4.0	80	1.70
5.0	2.5	4.0	80	2.09
6.0	2.5	4.0	80	2.17
7.0	2.5	4.0	80	2.45
5.0	1.5	4.0	80	2.62
5.0	2.0	4.0	80	2.42
5.0	3.0	4.0	80	1.01
5.0	3.5	4.0	80	0.82
5.0	2.5	2.0	80	0.59
5.0	2.5	6.0	80	3.20
5.0	2.5	8.0	80	4.50
5.0	2.5	10.0	80	7.81

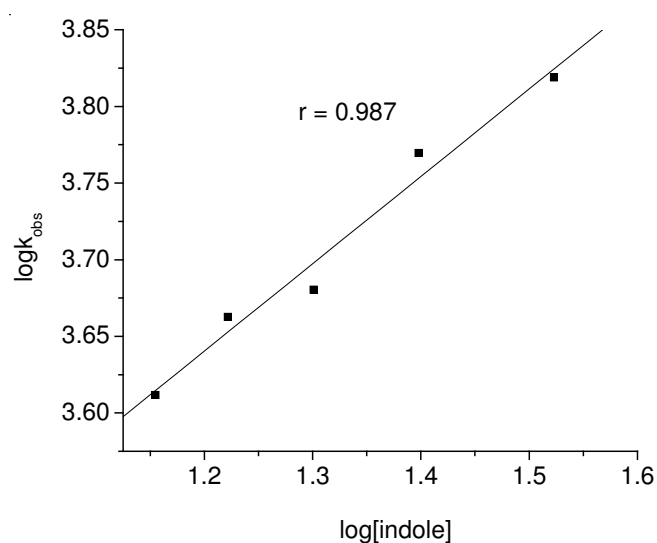


Fig. 1. Plot of  $\log k_{\text{obs}}$  *versus*  $\log$  [indole] for TPP catalyzed oxidation of indole by peroxomonosulphate in acetonitrile medium

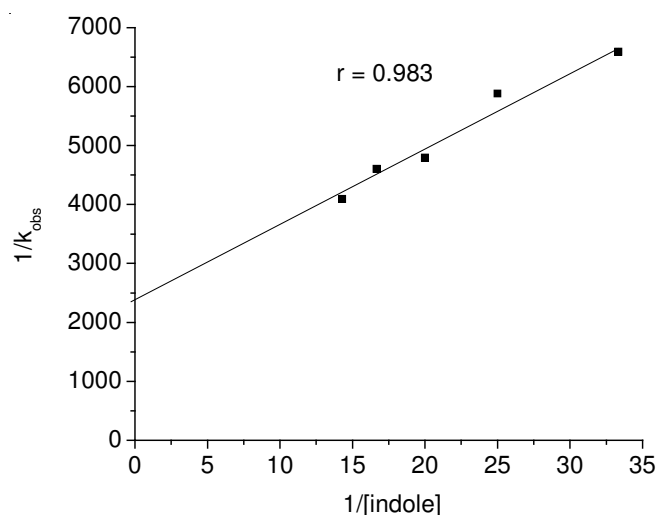


Fig. 2. Plot of  $1/k_{\text{obs}}$  *versus*  $1/[\text{indole}]$  for TPP catalyzed oxidation of indole by peroxomonosulphate in acetonitrile medium

with various initial concentrations of [TPP], which yielded rate constants whose values depended on [TPP]. The pseudo-first order rate constants ( $k_{\text{obs}}$ ) thus obtained were found to increase with [TPP] (Table-1) over a range of [TPP] used (0.2-

$1.0 \times 10^{-7} \text{ mol dm}^{-3}$ ). A linear plot was obtained between  $\log k_{\text{obs}}$  and  $\log [\text{TPP}]$  (Fig. 3) with a slope of 1.07, indicating that the order of the reaction with respect to catalyst was one.

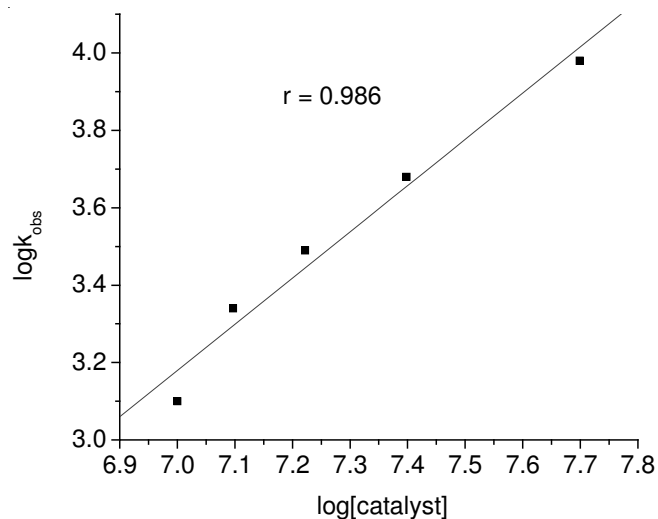


Fig. 3. Plot of  $\log k_{\text{obs}}$  versus  $\log [\text{catalyst}]$  showing the effect of catalyst concentration reaction rate

**Effect of  $[\text{H}^+]$ :** The reaction rates measured with various  $[\text{H}^+]$  ( $0.08\text{--}0.40 \text{ mol dm}^{-3}$ ) and were found to be the same (Table-2). Such kinetic behaviour indicates the nonexistence of any protonation equilibrium with respect to both PMS and indole in the experiment.

$[\text{H}^+] \times 10^2 (\text{mol dm}^{-3})$	$k_{\text{obs}} \times 10^4 (\text{s}^{-1})$
8.0	3.40
16.0	3.06
24.0	3.31
32.0	3.46
40.0	3.93

[PMS] =  $2.5 \times 10^{-3} \text{ mol dm}^{-3}$ , [indole] =  $5.0 \times 10^{-2} \text{ mol dm}^{-3}$ , [TPP] =  $4.0 \times 10^{-7} \text{ mol dm}^{-3}$ , acetonitrile:water = 40:10, temperature = 301 K.

**Effect of dielectric constant:** In order to determine the effect of dielectric constant (polarity) of the medium on rate, the oxidation of indole by peroxomonosulphate was studied in aqueous acetonitrile mixtures of various compositions (Table-3). The data clearly reveals that the rate increases with decrease in the percentage of acetonitrile, *i.e.*, with increasing dielectric constant or polarity of the medium and leads to the influence that there is a charge development in the transition state involving a more polar activated complex than the reactants<sup>20,21</sup>, a neutral molecule [indole] and a mono negative ion ( $\text{HSO}_5^-$ ) suggesting a polar (ionic) mechanism.

**Test for free radical intermediates:** No turbidity was observed when a freshly distilled acrylonitrile monomer was added to the deaerated reaction mixture indicating the absence of free radical intermediates.

**Rate and activation parameters:** The effect of temperature was studied in the range of 301–323 K and the results were shown in Table-4. The Arrhenius plot of  $\ln k_{\text{obs}}/T$  versus  $1/T$  was found to be linear (Fig. 4). The value of energy of activation ( $E_a$ ) was found to be  $10.36 \text{ kJ mol}^{-1} \text{ K}^{-1}$  and  $\Delta H^\ddagger =$

$\text{CH}_3\text{CN:H}_2\text{O}$	$D^a$	$k_{\text{obs}} \times 10^4 (\text{s}^{-1})$
90:10	41.47	0.31
80:20	45.48	2.09
70:30	49.71	2.76
60:40	53.83	5.42
50:50	57.95	8.95

[PMS] =  $2.5 \times 10^{-3} \text{ mol dm}^{-3}$ , [indole] =  $5.0 \times 10^{-2} \text{ mol dm}^{-3}$ , [TPP] =  $4.0 \times 10^{-7} \text{ mol dm}^{-3}$ . <sup>a</sup>Values are calculated from the values of pure solvent.

Temp. (K)	301	306	312	318	323
$k_{\text{obs}} \times 10^4 (\text{s}^{-1})$	1.86	2.09	3.34	4.66	5.73

[PMS] =  $2.5 \times 10^{-3} \text{ mol dm}^{-3}$ , [indole] =  $5.0 \times 10^{-2} \text{ mol dm}^{-3}$ , [TPP] =  $4.0 \times 10^{-7} \text{ mol dm}^{-3}$ . Acetonitrile:water = 40:10.

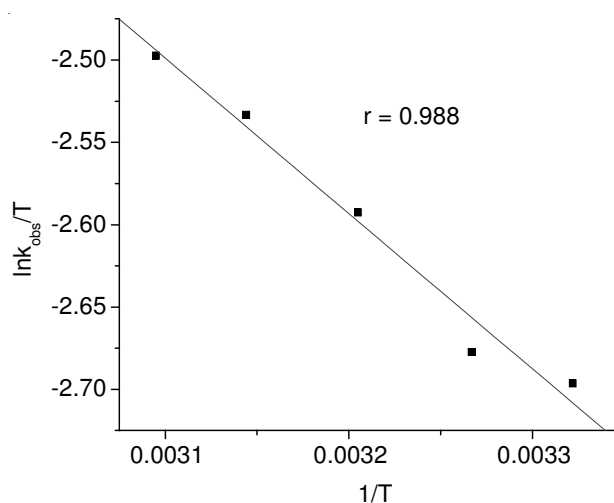


Fig. 4. Plot of  $\ln k_{\text{obs}}/T$  versus  $1/T$  showing the effect of temperature on reaction rate

$7.84 \text{ kJ mol}^{-1}$ ,  $\Delta S^\ddagger = -194.10 \text{ J K}^{-1} \text{ mol}^{-1}$ ,  $\Delta G^\ddagger = 66.65 \text{ kJ mol}^{-1}$ . The large negative value of entropy of activation ( $\Delta S^\ddagger$ ) obtained is attributed to the severe restriction of solvent molecules around the transition state<sup>22</sup>.

**Mechanism:** Peroxomonosulphate exists<sup>23</sup> as  $\text{HSO}_5^-$  in acidic solution. Although many peroxy anions are effective nucleophiles.  $\text{HSO}_5^-$  is weak nucleophile<sup>24</sup>. In spite of the fact that free radicals can arise from the facile homolysis of the oxygen-oxygen bond<sup>25</sup>, an ionic mechanism is favoured in certain reactions involving oxidations by peroxides. In the present investigation no observed polymerization in the presence of acrylonitrile rules out a free radical process. Hence ionic mechanism was favoured in this study. Generally, the enhancement of the electrophilic activity of peroxide will minimize the importance of undesirable free radical pathways, resulting in a mixture of products<sup>26</sup>. The Michaelis-Menton dependence of the oxidation rate on [indole] confirms the formation of complex in a rapid pre-equilibrium.

The first step is the formation of a complex between oxone and TPP which was confirmed from the UV-visible spectra (Fig. 5). This complex further decomposed and give  $\text{Por-Fe(V)} = \text{O}$  is in agreement with the literature study<sup>27</sup>. This  $\text{Por-Fe(V)} = \text{O}$  may further react with the indole to form a complex (Fig. 6)

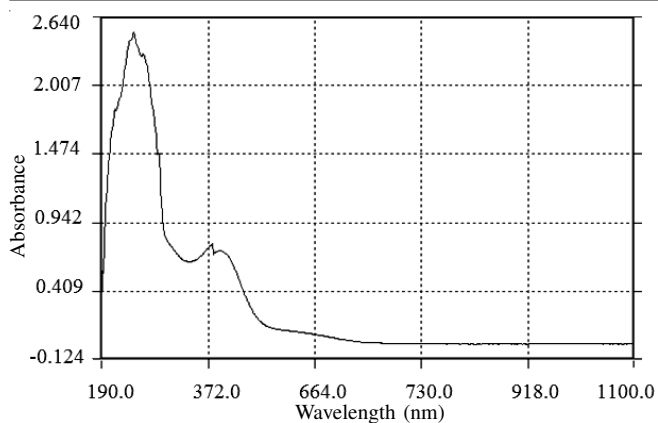


Fig. 5. UV spectrum showing the complex formation between TPP and PMS at 237 nm

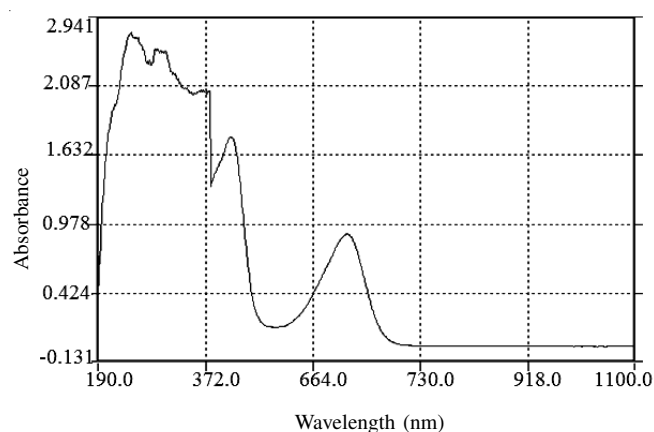


Fig. 6. UV spectrum showing formation of intermediate complex at 694 nm between indole and  $\text{PorFe(V)=O}$

at 694 nm, which would give the product in the next step (Fig. 7). This type of product was already reported<sup>9</sup>. The oxygen transfer step is associated with large negative value of entropies of activation and significant enthalpies of activation. The catalytic activity of TPP is significant and this conversion exhibits first order. The plot of  $\log [\text{catalyst}]$  versus  $\log k_{\text{obs}}$  is found to be linear. In accordance with the above observations and stoichiometry of the reaction, the following reactions are involved to constitute the most probable mechanism of the reaction (**Scheme-I**).

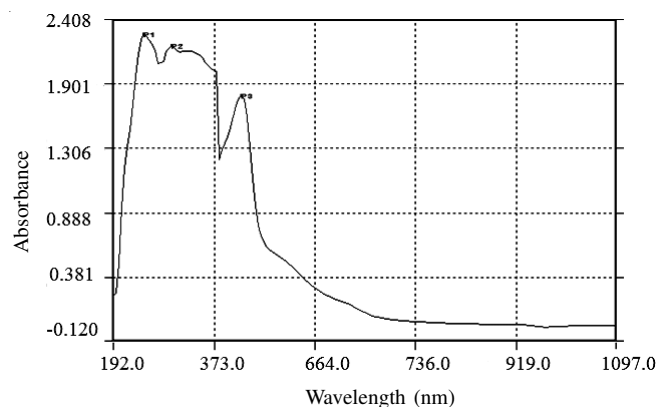
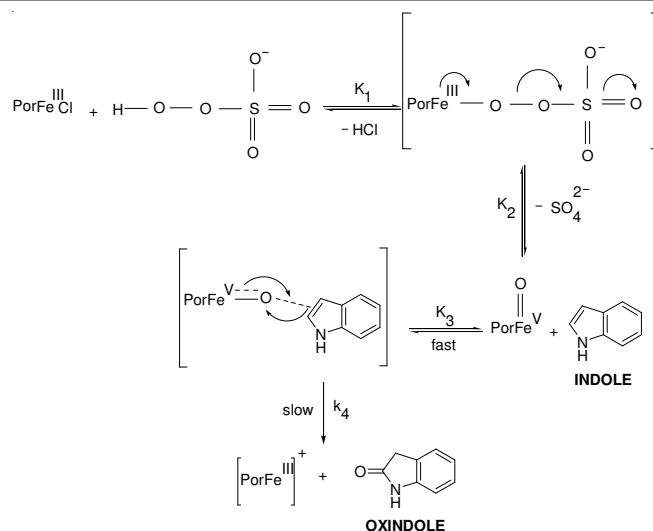


Fig. 7. UV spectrum showing the formation of product at 397 nm after the decomposition of intermediate complex at 694 nm



**Scheme-I:** Probable mechanism for the *meso*-tetraphenyl porphyriniron(III) chloride catalysed oxidation of indole by peroxomonosulphate

## REFERENCES

- R.J. Sundberg, Indole, Kirk-othmer Encyclopedia of Chemical Technology, Wiley, New York (1995).
- R.N. Goyal and A. Sangal, *Electrochim. Acta*, **50**, 2135 (2005).
- A.T. Al-kazwini, P. O'Neill, G.E. Adams, R.B. Cundall, G. Lang and A. Junino, *J. Chem. Soc. Perkin Trans II*, 1941 (1991).
- A.T. Al-kazwini, P. O'Neill, G.E. Adams, R.B. Cundall, A. Junino and J. Maignan, *J. Chem. Soc. Perkin Trans. II*, 657 (1992).
- S.N. Krylov and H.B. Dunford, *J. Phys. Chem.*, **100**, 913 (1996).
- W.B. Lawson and B. Witkop, *J. Org. Chem.*, **26**, 263 (1961).
- N. Finch and W.E. Taylor, *J. Am. Chem. Soc.*, **84**, 3871 (1962).
- K.S. Rangappa, D.T. Esterline, C.K. Mythily, D.S. Mahadevappa and S.Y. Ambedkar, *Polyhedron*, **12**, 1719 (1993).
- S.P. Meenakshisundaram and N. Sarathi, *Int. J. Chem. Kinet.*, **39**, 46 (2007).
- (a) R. Somthevan, R. Renganathan and P. Maruthamuthu, *Inorg. Chim. Acta*, **45**, L165 (1980); (b) A. Jameel and P. Maruthamuthu, *J. Indian Chem. Soc.*, **75**, 368 (1998).
- R. Suthakaran, P. Subramaniam and C. Srinivasan, *Proc. Indian Acad. Sci.*, **97**, 555 (1986).
- R.J. Kennedy and A.M. Stock, *J. Org. Chem.*, **25**, 1901 (1960).
- B.M. Trost and D.P. Curran, *Tetrahedron Lett.*, **22**, 1287 (1981).
- R. Mello, M. Fiorentino, C. Fusco and R. Curci, *J. Am. Chem. Soc.*, **111**, 6749 (1989).
- M. Balon, M. Munoz, P. Guardado, J. Hidalgo and C. Carmona, *Org. Chem.*, **58**, 7469 (1993).
- B. Meunier, *Chem. Rev.*, **92**, 1411 (1992).
- J. Larsen and K. Anker-Jorgensen, *J. Chem. Soc. Perkin Trans II*, 1213 (1992).
- J.T. Groves, T.E. Nemo and R.S. Myers, *J. Am. Chem. Soc.*, **101**, 1032 (1979).
- D.S. Bhuvaneshwari and K.P. Elango, *J. Indian Chem. Soc.*, **86**, 242 (2009).
- K.J. Laidler, Chemical Kinetics, Tata-McGraw Hill, New Delhi (1965).
- F. Ruff and A. Kucsman, *J. Chem. Soc. Perkin Trans. II*, 683 (1985).
- S.P. Meenakshisundaram, M. Selvaraju, N.M. Made Gowda and K.S. Rangappa, *Int. J. Chem. Kinet.*, **37**, 649 (2005).
- M.S. Ramachandran, T.S. Vivekanadam and V. Arunachalam, *Bull. Chem. Soc. (Japan)*, **59**, 1549 (1986).
- D.M. Davis and M.E. Deary, *J. Chem. Soc. Perkin Trans II*, 559 (1992).
- J.A. Kerr, *Chem. Rev.*, **66**, 465 (1966).
- Z. Zhu and J.H. Espenson, *J. Org. Chem.*, **60**, 1326 (1995).
- X.-T. Zhou, H.-B. Ji and Q.-L. Yuan, *J. Porphy. Phthalocya.*, **12**, 94 (2008).