

Synthesis of Nitriles from Aldehydes and Hydroxylamine Hydrochloride using Anhydrous Sodium Sulphate and Sodium Bicarbonate in Dry Media

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A rapid and facile one pot synthesis of nitriles has been carried out in high yields from the corresponding aldehydes and hydroxylamine hydrochloride in the presence of environmentally benign anhydrous sodium sulphate and sodium bicarbonate catalysts in dry media under microwave irradiation.

Key Words: Nitriles, Aldehydes, Hydroxylamine hydrochloride, Anhydrous sodium sulphate, Sodium bicarbonate.

INTRODUCTION

Nitriles are widely used for transformation into amides, amines, esters, carboxylic acids¹ *etc.* Hence they have been used as intermediates for the synthesis of fine chemicals such as agricultural chemicals, dyes and medicines². One of the most general methods for the synthesis of nitriles is the nucleophilic substitution reaction of alkyl halides with metal cyanides. The method is, however, inconvenient because of high toxicity of metal cyanides and troublesome handling. Consequently, other methods such as the dehydration of primary amides³ or aldoximes⁴⁻⁷ have attracted attention. It is known that dehydration of aldoximes into nitriles can be achieved by using a variety of reagents like triethylamine/sulfur dioxide⁴, zeolites⁵, sulfurylchloride fluoride⁶, sulfuryl chloride⁷, *etc.* but many of these suffer from limitations such as high toxicity, vigorous reaction conditions, unsatisfactory yields, tedious work up and use of large excess of reagents. Recently, a rapid synthesis of nitriles in high yields from aldoximes using silica gel^{8a} and one-pot

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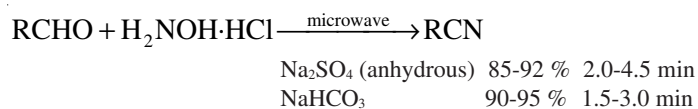
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synthesis of nitriles from aldehydes and hydroxylamine hydrochloride using silicagel, Mont K-10 and KSF catalysts in dry media under microwave irradiation have been reported⁹. Some rapid procedures for one pot synthesis of nitriles have been described using formic acid¹⁰ and potassium peroxy monosulfate¹¹ but whereas the first method suffers from undesirable action of formic acid that can affect acid-sensitive aldehydes, the second suffers from the undesirable oxidation of some functional groups. So, there exists a need for developing rapid and facile methods for one pot synthesis of nitriles. As part of our interest in rapid synthetic transformations¹²⁻¹⁴. Herein a rapid one pot synthesis of nitriles aldehydes and hydroxylamine hydrochloride using anhydrous sodium sulphate and sodium bicarbonate in dry media are reported, which would catalyze both the oxime formation and the consequent nitrile formation steps.

EXPERIMENTAL

Melting points were determined in open capillaries on an electrically heated metal block and are uncorrected. ¹H NMR (CDCl₃) spectra were recorded on a Jeol FX90Q instrument using TMS as an internal standard. IR spectra were recorded on a Perkin-Elmer 782 spectrophotometer. TLC was carried out on silica gel G plates with a benzene/ethyl acetate (4:1) system.

Synthesis: 3,4-Dimethoxybenzaldehyde (166 mg, 1 mmol), hydroxylamine hydrochloride (105 mg, 1.2 mmol) were mixed with anhydrous sodium sulphate (1 g) or (NaHCO₃, 1 g) and taken up in an Erlenmeyer flask (25 mL) and irradiated at 560 W for 4 min in an unmodified domestic microwave oven (Kenstar OM-9925E, 800 W, operating at 2450 MHz). The flask was taken out, cooled and ether (15 mL) was added. The catalyst was filtered off and the resultant solution evaporated to give a residue which was purified by chromatography using benzene:ethyl acetate (4:1) as eluent to afford the desired nitrile (85 % yield with Na₂SO₄ and 93 % yield with NaHCO₃) (**Scheme-I**): ¹H NMR (CDCl₃), 3.8 (s, 3H, -OCH₃), 3.9 (s, 3H, -OCH₃) 6.9 (m, 1H, ArH), 7.0 (m, 2H, ArH); IR (neat) 2940 ν(C-H *str.*) 2220 ν(C≡N *str.*), 1610, 1540 and 1450 ν(C-C *str.*); m.p. 62.5 °C, lit⁸ 63 °C.



Scheme-I

RESULTS AND DISCUSSION

Microwave activation as a non-conventional energy source especially when coupled to dry-media conditions has become a very popular and useful technology in organic chemistry in terms of reduced reaction times, enhanced product purity and chemical yields in an environmentally friendly approach called Green Chemistry¹⁵. Dry media conditions are particularly suitable for doing away with the safety hazards usually associated with low-boiling solvents under microwave irradiation.

Investigations were initiated with 3,4-dimethoxybenzaldehyde being chosen as a model compound. It was condensed with hydroxylamine hydrochloride in presence of sodium sulphate (anhyd) under varying levels of irradiation (80-800 W). The best yield of 85 %, was obtained at 560 W after 4.5 min of irradiation. Hence, all the subsequent reactions were carried out under these reaction conditions. A variety of substituted aromatic aldehydes (**1-6**), quinoline carboxaldehyde (**8**), pyridine carboxaldehyde (**9**) and an aliphatic aldehyde (**10**) were included. The reactions were monitored by TLC. The products yields and reactions times are given in the Table-1. All the products are known compounds and were identified on the basis of their spectroscopic analyses and by direct comparison of their m.p. with those of the authentic samples^{8,9}. The products were obtained in more than 95 % purity as determined by ¹H NMR and were further purified by chromatography.

Next, one pot synthesis of nitriles using NaHCO₃ as the catalyst were also carried out. The nitriles were obtained in slightly better yields in 90-

TABLE-1
ONE POT SYNTHESIS OF NITRILES FROM ALDEHYDES AND
HYDROXYLAMINE HYDROCHLORIDE USING ANHYDROUS
SODIUM SULPHATE AND SODIUM BICARBONATE AT 560 W

Entry	Product	Na ₂ SO ₄ (anhyd)		NaHCO ₃	
		Isolated yield (%)	Time (min)	Isolated yield (%)	Time (min)
1	3,4-Dimethoxy-benzonitrile	85	4.5	93	3.0
2	4-Methoxybenzonitrile	92	3.5	95	1.5
3	2-Hydroxybenzonitrile	90	3.5	90	1.5
4	4-Hydroxybenzonitrile	89	4.0	90	1.5
5	4-Nitrobenzonitrile	86	2.0	94	1.5
6	3,4,5-Trimethoxy benzonitrile	89	3.0	90	2.5
7	<i>trans</i> -Cinnamitrile	90	2.0	91	2.0
8	Quinoline-2-carbonitrile	89	2.0	91	1.5
9	2-Ethyl-4-cyanopyridine	90	3.0	93	2.5
10	Octanenitrile	90	3.0	90	2.5

95 % under identical reaction conditions as compared to those obtained with anhydrous sodium sulphate (Table-1). It should be noted that the nitriles were formed in only 2-3 % yield under identical reaction conditions in the absence of the catalysts, thereby confirming the role of the catalysts in the reactions.

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